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materials, important in the development of new materials and the basis of many biological systems, cannot be designed using trial and error methods due to the multiplicity of components and parameters.

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Of Surfaces And
Interfaces Omid Moradi
Shahre-Qods Branch,
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Introduction

Thermodynamics is the branch of science that is concerned with the principles of energy transformation in macroscopic systems. Macroscopic properties of matter arise from the behavior of a very large number of molecules.

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Understanding the
structural and
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properties of surfaces,
interfaces, and
membranes is
important for both
fundamental and
practical reasons.
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include coatings,

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dispersants,
encapsulating agents,
and biological
materials.

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Summary of basic thermodynamic concepts (Chapter 1

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dealing with?

Thermodynamics is the
branch of science that
is concerned with the
principles of energy
transformation in
macroscopic system.
Macroscopic properties
of matter arise from
the behavior of a very
large number of
molecules.

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summary of
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principles that are
relevant to interfaces
in view of the topics
discussed such as

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systems, Equilibrium
between phases,
Physical description of
a real liquid interface,
Surface free energy
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liquids, Surface
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Relation of van der
Waals constants with
molecular pair
potentials and etc in
forthcoming and
special attention is
paid to heterogeneous

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systems that contain
phase ...

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different from atoms in
the bulk. • The fewer
neighbors of the
surface cause it to
have a very different
and anisotropic
chemical environment
compared with the

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bulk. • The thermodynamics of the surface is most likely to be quite different from the thermodynamic properties of the bulk.

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Introduction to surface quantities (Chapter 2, ...

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Unanticipated nanostructures, characterized by the presence of phases at interfaces and surfaces which are unstable as bulk phases, can be thermodynamically stabilized due to the dominance of energy contributions of interfaces and surfaces in the total Gibbs energy of the system.

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An interface constitutes the separation surface between two phases. A phase constitutes a homogeneous part of a thermodynamic system, i.e., the part of the universe that is under investigation. The system has always to be defined by the investigator before she/he starts to perform some studies.

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From Surface
Science,
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In a wider physicochemical perspective, Equation designates the “surface law of mass action” of the system that suggests that Le Châtelier's principle—if a chemical system at equilibrium experiences a change in temperature, pressure, and species concentration, then the equilibrium shifts to counteract the imposed change and a

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new equilibrium is established—must be complemented at interfaces by a term accounting for the surface work.

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surface phenomena are detailed, including wetting, crystalline systems (including grain boundaries), interfaces between different phases,

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curved interfaces (capillarity), adsorption phenomena and adhesion of surface layers. The later chapters also feature case studies to illustrate real-world applications.

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