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For the levels to differ the pressure P_1 must be greater than P_2 , hence. $P_1 = P_2 + h\rho g$. If P_1 is the lung pressure, P_0 is the atmospheric pressure, then if the difference is 'h' then lung pressure can be calculated as follows. $P_1 = P_0 + h\rho g$. Example. A man blows into one end of a U-tube containing water until the levels differ by 40.0 cm. if the atmospheric pressure is 1.01×10^5 ...

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Let the pressure due to water be $h_1 \rho_1 g_1$ and that of water be $h \rho g$. Then $h_1 \rho_1 g_1 = h \rho g$. Hence $h_1 = (6.7 \times 10^{-1}) \times (1.36 \times 10^4) / 1.0 \times 10^3 = 911.2$ cm or 9.11 m.

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